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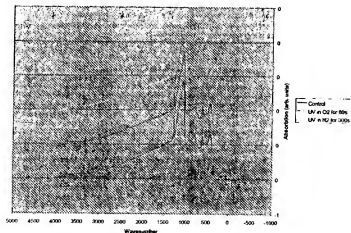
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(54) Title: ULTRAVIOLET CURING PROCESS FOR POROUS LOW-K MATERIALS



(57) Abstract: Low dielectric constant porous materials with improved elastic modulus. The process of making such porous materials involves providing a porous dielectric material and ultraviolet (UV) curing of the porous dielectric material to produce a UV cured porous dielectric material. UV curing of the porous dielectric material yields a material with improved modulus and comparable dielectric constant. The improvement in elastic modulus is typically greater than about 50%. The porous dielectric material is UV cured for no more than about 300 seconds at a temperature less than about 450°C. The UV cured porous dielectric material can optionally be post-UV treated. Rapid Anneal Processing (RAP) of the UV cured porous dielectric material reduces the dielectric constant of the material while maintaining an improved elastic modulus as compared to the UV cured porous dielectric material. The annealing temperature is typically less than about 450°C, and the annealing time is typically less than about 60 minutes. The post-UV treated, UV cured porous dielectric material has a dielectric constant between about 1.1 and about 3.5 and an improved elastic modulus.

## ULTRAVIOLET CURING PROCESS FOR POROUS LOW-K MATERIALS

### CROSS-REFERENCE TO RELATED APPLICATIONS

- This application is related to U.S. Patent Application Serial No. 09/528,835, filed March 20, 2000 and entitled "HIGH MODULUS, LOW DIELECTRIC CONSTANT COATINGS" and U.S. Patent Application Serial No. 09/681,332, filed March 19, 2001 and entitled "PLASMA CURING PROCESS FOR POROUS SILICA THIN FILM", the disclosures of which are incorporated herein by reference.

### BACKGROUND OF THE INVENTION

This invention relates generally to a process which is employed in manufacturing semiconductor chips. More particularly, the invention relates to a process for improving the structural properties of certain porous materials that are utilized as integrated circuit (IC) dielectrics.

- New materials with low dielectric constants (known in the art as "low-k dielectrics") are being investigated for their potential use as insulators in semiconductor chip designs. A low dielectric constant material aids in enabling further reductions in the integrated circuit feature dimensions. The substance with the lowest dielectric constant is air ( $k=1.0$ ). Therefore, porous dielectrics are very promising candidates since they have the potential to provide very low dielectric constants. Unfortunately, however, such porous low-k dielectrics typically have the problem of insufficient mechanical strength.

- Thin film dielectric coatings on electric devices are known in the art. For instance, U.S. Patent Nos. 4,749,631 and 4,756,977, to Haluska et al., disclose silica based coatings produced by applying solutions of silicon alkoxides or hydrogen silsesquioxane, respectively, to substrates and then heating the coated substrates to a temperature between 200 and 1000°C. The dielectric constant of these coatings is often too high for certain electronic devices and circuits.

- U.S. Patent Nos. 4,847,162 and 4,842,888, to Haluska et al., teach the formation of nitrified silica coatings by heating hydrogen silsesquioxane resin and silicate esters, respectively, to a temperature between 200 and 1000°C in the presence of ammonia.

Glasser et al., Journal of Non-Crystalline Solids, 64 (1984) pp. 209-221, teaches the formation of ceramic coatings by heating tetraethoxysilane in the

presence of ammonia. This reference teaches the use of anhydrous ammonia and that the resulting silica coatings are nitrified.

5 U.S. Patent No. 4,636,440, to Jada, discloses a method of reducing the drying time for a sol-gel coated substrate comprising exposing the substrate to aqueous quaternary ammonium hydroxide and/or alkanol amine compounds. Jada requires that the coating be dried prior to heating. It is specifically limited to hydrolyzed or partially hydrolyzed silicon alkoxides.

10 U.S. Patent Nos. 5,262,201, to Chandra, and 5,116,637, to Baney et al., teach the use of basic catalysts to lower the temperature necessary for the conversion of various preceramic materials, all involving hydrogen silsesquioxane, to ceramic coatings. These references teach the removal of solvent before the coating is exposed to the basic catalysts.

15 U.S. Patent No. 5,547,703, to Camilletti et al., teaches a method for forming low dielectric constant Si-O containing coatings on substrates comprising heating a hydrogen silsesquioxane resin successively under wet ammonia, dry ammonia, and oxygen. The resultant coatings have dielectric constants as low as 2.42 at 1 MHz. This reference teaches the removal of solvent before converting the coating to a ceramic.

20 U.S. Patent No. 5,523,163, to Balance et al., teaches a method for forming Si-O containing coatings on substrates comprising heating a hydrogen silsesquioxane resin to convert it to a Si-O containing ceramic coating and then exposing the coating to an annealing atmosphere containing hydrogen gas. The resultant coatings have dielectric constants as low as 2.773. The reference teaches the removal of solvent before converting the coating to a ceramic.

25 U.S. Patent No. 5,618,878, to Syktich et al., discloses coating compositions containing hydrogen silsesquioxane resin dissolved in saturated alkyl hydrocarbons useful for forming thick ceramic coatings. The alkyl hydrocarbons disclosed are those up to dodecane. The reference does not teach exposure of the coated substrates to basic catalysts before solvent removal.

30 U.S. Patent Application Serial No. 09/197,249, to Chung et al., entitled "A METHOD OF FORMING COATINGS" and filed November 20, 1998, discloses a method of making porous network coatings with low dielectric constants. The method comprises depositing a coating on a substrate with a solution comprising a resin containing at least 2 Si-H groups and a solvent in a manner in which at least 5

volume % of the solvent remains in the coating after deposition. The coating is then exposed to an environment comprising a basic catalyst and water. Finally, the solvent is evaporated from the coating to form a porous network. If desired, the coating can be cured by heating to form a ceramic. Films made by this process have dielectric constants in the range of 1.5 to 2.4 with an elastic modulus between about 2 and about 3 GPa.

As was described in U.S. Patent Application Serial No. 09/681,332, entitled "PLASMA CURING PROCESS FOR POROUS SILICA THIN FILM" and incorporated herein by reference, instead of plasma treating, porous network coatings produced from a resin containing at least 2 Si-H groups can be plasma cured, eliminating the need for prior furnace curing.

However, there remains a need for a process for making other porous low-k material with improved structural properties, such as an improved elastic modulus, without compromising or deteriorating its electrical properties.

#### SUMMARY OF THE INVENTION

The present invention meets that need by providing a process that produces materials having a low dielectric constant and an improved elastic modulus. The process involves providing a porous dielectric material having a first dielectric constant and having a first elastic modulus. The porous dielectric material is cured with ultraviolet (UV) radiation to produce a UV cured porous dielectric material having a second dielectric constant which is comparable to the first dielectric constant and having a second elastic modulus which is greater than the first elastic modulus. This increase in elastic modulus is typically greater than about 50%.

The UV cured porous dielectric material can optionally be post-UV treated to provide a post-UV treated, UV cured porous dielectric material having a third dielectric constant and having a third elastic modulus. Post-UV treatment of the UV cured porous dielectric material reduces the dielectric constant of the porous dielectric material while maintaining the increase in the elastic modulus as compared to the elastic modulus before the post-UV treatment.

Accordingly, it is an object of the present invention to produce porous dielectric materials having an improved elastic modulus and a low dielectric constant.

#### BRIEF DESCRIPTION OF THE DRAWINGS

Fig. 1 is a chart illustrating the FTIR spectra for MSQ-based porous low-k films which were thermally cured, thermally cured and UV cured in O<sub>2</sub>, and thermally cured and UV cured in N<sub>2</sub>.

Fig. 2 is a chart illustrating the FTIR spectra for HSQ-based porous low-k films which were uncured, UV cured in O<sub>2</sub>, and UV cured in N<sub>2</sub>.

Fig. 3 is a chart illustrating the FTIR spectra for 5% MSQ/95% HSQ-based porous low-k films which were uncured, UV cured in O<sub>2</sub>, and UV cured in N<sub>2</sub>.

#### DETAILED DESCRIPTION OF THE INVENTION

The present invention is based on the discovery that UV curing virtually any porous dielectric material, without the necessity of thermally curing the material, increases the elastic modulus (Young's modulus) and material hardness of the porous material while maintaining its low dielectric constant properties. The porous dielectric material can include, but is not limited to, hydrogen silsesquioxane (HSQ) dielectric materials, methylsilsesquioxane (MSQ) dielectric materials, organic dielectric materials, inorganic dielectric materials, and combinations thereof, which can be produced by spin-on or chemical vapor deposition (CVD) processes. The porous dielectric materials can have porogen-generated, solvent-based, or molecular engineered pores, which may be interconnected or closed, and which may be distributed random or ordered, such as vertical pores.

UV curing can generate a notable amount of polar species in the porous dielectric material, which can be undesirable in some applications. The present invention is also based on the discovery that post-UV treating UV cured porous dielectric materials produces a low dielectric constant, improved modulus material.

The process of the present invention is particularly applicable to the deposition of coatings on electronic devices or electronic circuits where they can serve as interlevel dielectric layers, doped dielectric layers to produce transistor-like devices, pigment loaded binder systems containing silicon to produce capacitor and capacitor-like devices, multilayer devices, 3-D devices, silicon on insulator devices, super lattice devices, and the like. However, the choice of substrates and devices to be coated by the instant invention is limited only by the need for thermal and chemical stability of the substrate at the temperature and pressure used in the present invention. As such, the porous dielectric materials of the present invention can be used on substrates such as plastics including, for example, polyimides,

epoxies, polytetrafluoroethylene and copolymers thereof, polycarbonates, acrylics and polyesters, ceramics, leather, textiles, metals, and the like.

As used in the present invention, the expression "ceramic" includes ceramics such as amorphous silica and ceramic-like materials such as amorphous silica-like materials that are not fully free of carbon and/or hydrogen but are otherwise ceramic in character. The expressions "electronic device" or "electronic circuit" include, but are not limited to, silica-based devices, gallium arsenide based devices, silicon carbide based devices, focal plane arrays, opto-electronic devices, photovoltaic cells, and optical devices.

A porous dielectric material is needed as a starting material for the present invention. Typical HSQ-based dielectric materials for use with the present invention include FOx HSQ-based dielectric material and XLK porous HSQ-based dielectric material available from Dow Corning Corporation (Midland, MI). In addition, typical ultra low-k porous dielectric MSQ-based materials, made by spin-on processing, for use with the present invention are available from Chemat Technology, Inc. (Northridge, CA) and JSR Corporation (Tokyo, Japan).

The production of typical porous dielectric materials for use with the present invention is well known in the art. One method of making such a porous dielectric material is the porous network coating disclosed in U.S. Patent Application Serial No. 09/197,249, which is incorporated herein by reference for its teaching on how to produce porous dielectric materials having ultra low dielectric constants. The application describes the manufacture of ultra low dielectric constant coatings having a dielectric constant between about 1.5 and about 2.4, in which pores are introduced into HSQ-based films. HSQ-based films produced according to the method taught in U.S. Patent Application Serial No. 09/197,249, which have been cured under thermal conditions, contain about 20 to about 60% Si-H bonds density. When the dielectric constant of the coating is about 2.0, the coating has an elastic modulus of between about 2 and about 3 GPa.

The following method of producing a porous network coating is provided as an example of the production of a typical porous dielectric material. It is not the inventors' intent to limit their invention to only HSQ-based films. The process of the present invention is applicable to virtually any porous dielectric material.

The method of producing the HSQ-based porous network coating starts with depositing a coating on a substrate with a solution comprising a resin containing at

least 2 Si-H groups and a solvent. The resins containing at least 2 Si-H groups are not particularly limited, as long as the Si-H bonds can be hydrolyzed and at least partially condensed by the basic catalyst and water to form a cross-linked network that serves as the structure for the porous network. Generally, such materials have

5 the formula:



wherein each R is independently selected from the group consisting of hydrogen, alkyl, alkenyl, or aryl groups, or alkyl, alkenyl, or aryl groups substituted with a hetero atom such as a halogen, nitrogen, sulfur, oxygen, or silicon, and a, b, c, and d are mole fractions of the particular unit and their total is 1, with the proviso that at least 2 R groups per molecule are hydrogen and the material is sufficiently resinous in structure to form the desired network. Examples of alkyl groups are methyl, ethyl, propyl, butyl, and the like, with alkyls of 1-6 carbons being typical. Examples of alkenyl groups include vinyl, allyl, and hexenyl. Examples of aryls include phenyl.

15 Examples of substituted groups include  $CF_3(CF_2)_nCH_2CH_3$ , where  $n = 0-6$ .

Useful in the present invention are various hydrosiloxane resins, known as hydrogen silsesquioxane resins, comprising units of the formula  $HSi(OH)_x(OR')_yO_{2z/2}$ . In this formula, each R' is independently selected from the group consisting of alkyl, alkenyl, or aryl groups, or alkyl, alkenyl, or aryl groups substituted with a hetero atom such as a halogen, nitrogen, sulfur, oxygen, or silicon. Examples of alkyl groups are methyl, ethyl, propyl, butyl, and the like, with alkyls of 1-6 carbons being typical. Examples of alkenyl groups include vinyl, allyl, and hexenyl. Examples of aryls include phenyl. Examples of substituted groups include  $CF_3(CF_2)_nCH_2CH_3$ , where  $n = 0-6$ . When these R' groups are bonded to silicon through the oxygen atom, they form a hydrolyzable substituent. In the above formula,  $x = 0$  to 2,  $y = 0$  to 2,  $z = 1$  to 3, and  $x + y + z = 3$ . These resins may be essentially fully condensed ( $HSiO_{3/2}$ ) $_n$  where  $n$  is 8 or greater, or they may be only partially hydrolyzed (i.e., containing some Si-OR'), and/or partially condensed (i.e., containing some Si-OH).

The structure of the resin containing at least 2 Si-H groups is not limited. The structure may be what is generally known as ladder-type, cage-type, or mixtures thereof. The HSQ resins may contain endgroups such as hydroxyl groups, triorganosiloxy groups, diorganohydrogensiloxy groups, trialkoxy groups, dialkoxy groups, and others. The HSQ resin may also contain a small number (e.g., less than

10%) of the silicon atoms, which have either 0 or 2 hydrogen atoms attached thereto and/or a small number of Si-C groups, such as  $\text{CH}_3\text{SiO}_{3/2}$  or  $\text{HCH}_3\text{SiO}_{2/2}$  groups.

The resins containing at least 2 Si-H groups and methods for their production are known in the art. For example, U.S. Patent No. 3,615,272, to Collins, teaches  
5 the production of an essentially fully condensed hydrogen silsesquioxane resin (which may contain up to 100-300 ppm silanol) by a process comprising hydrolyzing trichlorosilane in a benzenesulfonic acid hydrate hydrolysis medium, and then washing the resulting resin with water or aqueous sulfuric acid. Similarly, U.S. Patent No. 5,010,159, to Bank, teaches a method comprising hydrolyzing  
10 hydridosilanes in an arylsulfonic acid hydrate hydrolysis medium to form a resin which is then contacted with a neutralizing agent.

Other hydridosiloxane resins, such as those described in U.S. Patent No. 4,999,397, to Frye, and U.S. Patent No. 5,210,160, to Bergstrom, those produced by hydrolyzing an alkoxy or acyloxy silane in an acidic, alcoholic hydrolysis medium,  
15 those described in Japanese Kokai Patent Nos. 59-178749, 60-86017, and 63-107122, or any other equivalent hydridosiloxanes, will also function herein.

Specific molecular weight fractions of the Si-H containing resins may also be used. Such fractions and methods for their preparation are taught in U.S. Patent No. 5,063,267, to Hanneman, and U.S. Patent No. 5,416,190, to Mine. A typical fraction  
20 comprises material wherein at least 75% of the polymeric species have a number average molecular weight above about 1200, and a more typical fraction comprises material wherein at least 75% of the polymeric species have a number average molecular weight between about 1200 and about 100,000.

The Si-H containing resins may contain other components as long as these  
25 components do not interfere with the integrity of the coating. It should be noted, however, that certain materials may increase the dielectric constant of the coating.

Ceramic oxide precursors may also be used in combination with the Si-H containing resins. The ceramic oxide precursors useful herein include compounds of various metals such as aluminum, titanium, zirconium, tantalum, niobium and/or  
30 vanadium, as well as various non-metallic compounds, such as those of boron or phosphorus, which may be dissolved in solution, hydrolyzed and subsequently pyrolyzed at relatively low temperature to form ceramic oxides. Ceramic oxide precursors useful herein are described in U.S. Patent Nos. 4,808,653, 5,008,320, and 5,290,394.



The Si-H containing resins are applied to the substrates as solvent dispersions to form a coating on the substrate ("SiH resin coating"). Solvents that may be used include any agent or mixture of agents that will dissolve or disperse the resin to form a homogeneous liquid mixture without affecting the resulting coating or the substrate. These solvents can include alcohols, such as ethyl alcohol or isopropyl alcohol; aromatic hydrocarbons, such as benzene or toluene; branched or linear alkanes, such as n-heptane, dodecane, or nonane; branched or linear alkenes, such as n-heptene, dodecene, or tetradecene; ketones, such as methyl isobutyl ketone; esters; ethers, such as glycol ethers; or linear or cyclic siloxanes, such as hexamethyldisiloxane, octamethyldisiloxane, and mixtures thereof, or cyclic dimethylpolysiloxanes; or mixtures of any of the above solvents. The solvent is generally present in an amount sufficient to dissolve/disperse the resin to the concentration desired for application. Typically, the solvent is present in an amount of about 20 to about 99.9 wt%, and more typically from about 70 to about 95 wt%, based on the weight of the resin and solvent.

If desired, other materials can be included in the resin dispersion. For instance, the dispersion can include fillers, colorants, adhesion promoters, and the like.

Specific methods for application of the resin dispersion to the substrate include, but are not limited to, spin coating, dip coating, spray coating, flow coating, screen printing, or others. A typical method is spin coating.

At least about 5 volume % of the solvent should remain in the SiH resin coating until the resin is contacted with the basic catalyst and water. This solvent forms the pores of the porous network coating as the Si-H bonds are hydrolyzed and condensed. In some embodiments, it may be typical that at least about 10 volume % solvent remains, while in others, it may be typical that at least about 15 volume % solvent remains, and in still others, it may be typical that at least about 25 volume % solvent remains.

The method of retaining the solvent is not particularly restricted. In a typical embodiment, a high boiling point solvent can be used alone or as a co-solvent with one of the solvents described above. In this manner, processing the resin dispersion as described above under normal conditions allows for at least about 5% residual solvent remaining. Typical high boiling solvents in this embodiment are those with boiling points above about 175°C including hydrocarbons, aromatic hydrocarbons,

esters, ethers, and the like. Examples of specific solvents which can be used in this embodiment include saturated hydrocarbons, such as dodecane, tetradecane, hexadecane, etc., unsaturated hydrocarbons, such as dodecene, tetradecene, etc., xylenes, mesitylene, 1-heptanol, dipentene, d-limonene, tetrahydrofurfuryl alcohol, mineral spirits, 2-octanol, stoddard solvent, Isopar H™, diethyl oxalate, diamyl ether, tetrahydropyran-2-methanol, lactic acid butyl ester, isooctyl alcohol, propylene glycol, dipropylene glycol monomethyl ether, diethylene glycol diethyl ether, dimethyl sulfoxide, 2,5-hexanedione, 2-butoxyethanol acetate, diethylene glycol monomethyl ether, 1-octanol, ethylene glycol, Isopar L™, dipropylene glycol monomethyl ether acetate, diethylene glycol monoethyl ether, N-methylpyrrolidone, ethylene glycol dibutyl ether, gamma-butyrolactone, 1,3-butanediol, diethylene glycol monomethyl ether acetate, trimethylene glycol, triethylene glycol dimethyl ether, diethylene glycol monoethyl ether acetate, alpha-terpineol, n-hexyl ether, kerosene, 2-(2-n-butoxyethoxy)ethanol, dibutyl oxalate, propylene carbonate, propylene glycol monophenyl ether, diethylene glycol, catechol, diethylene glycol monobutyl ether acetate, ethylene glycol monophenyl ether, diethylene glycol dibutyl ether, diphenyl ether, ethylene glycol monobenzyl ether, hydroquinone, sulfolane, and triethylene glycol. Hydrocarbon solvents are particularly preferred.

The above processing (i.e., primarily deposition of the SiH resin coating solution) can be done in an environment that inhibits solvent evaporation prior to contact with the basic catalyst and water. For example, the spin coating can be performed in a closed environment such that the subsequent steps (i.e., contact with the basic catalyst and water) can occur before the solvent is completely evaporated.

The SiH resin coating containing at least about 5 volume % solvent is then contacted with a basic catalyst and water. Examples of basic catalysts include ammonia, ammonium hydroxide, as well as amines. The amines useful herein may include primary amines ( $RNH_2$ ), secondary amines ( $R_2NH$ ), and/or tertiary amines ( $R_3N$ ) in which R is independently a saturated or unsaturated aliphatic, such as methyl, ethyl, propyl, vinyl, allyl, ethynyl, etc.; an alicyclic, such as cyclohexylmethyl; an aromatic, such as phenyl; a substituted hetero atom, such as oxygen, nitrogen, sulfur, etc.; or compounds in which the nitrogen atom is a member of a heterocyclic ring such as quinoline, pyrrolidine, or pyridine. In addition, any of the above amine compounds may be substituted with other hydrocarbon and/or hetero containing groups to form compounds such as diamines, amides, etc. Finally, it is also

contemplated that compounds, which are converted to amines under the reactions conditions used, would function in an equivalent manner. For example, a compound such as an ammonium salt that yields an amine upon dissolution would provide the desired catalytic effect.

- 5           Examples of the amines that may be used herein include methylamine, ethylamine, butylamine, allylamine, cyclohexylamine, aniline, dimethylamine, diethylamide, dioctylamine, dibutylamine, methylethylamine, saccharin, piperidine, trimethylamine, triethylamine, pyridine, diethyl toluidene ethylmethylpropylamine, imidazole, choline acetate, triphenyl phosphine aniline, trimethylsilylimidazole, ethylenediamine, diethylhydroxylamine, triethylenediamine, n-methylpyrrolidone, etc.

10           The basic catalyst can generally be used at any concentration sufficient to catalyze hydrolysis of the Si-H bonds. Generally, concentrations of the basic catalyst can be from about 1 ppm to about 100 wt% based on the weight of the resin, depending on the basic catalyst.

- 15           The water used can be that present in the ambient environment (e.g., > about 25% relative humidity), the ambient environment can be supplemented with additional water vapor (e.g., relative humidity up to about 100%), water can be used as a liquid, or a compound which generates water under the reaction conditions can be used.

- 20           Contact of the SiH resin coating with the basic catalyst and water can be accomplished by any means practical or desirable. For instance, the SiH resin coating can be contacted with vapors of the basic catalyst and water vapor. Alternatively, the SiH resin coating can be contacted with the basic catalyst and water in the liquid state, such as by immersing the coating in an ammonium  
25           hydroxide solution.

- The SiH resin coating is typically exposed to an environment comprising the basic catalyst and water in the vapor state, more typically ammonia and water vapor. For instance, the SiH resin coated substrate may be placed in a container and the appropriate environment introduced therein, or a stream of the basic catalyst and  
30           water may be directed at the SiH resin coating.

          The method used to generate the basic catalyst and water environment is generally not significant in the present embodiment. Methods such as bubbling the basic catalyst (e.g., ammonia gas) through water or ammonium hydroxide solutions (to control the amount of water vapor present), heating a basic catalyst and water, or

heating water and introducing the basic catalyst gas (e.g., ammonia gas) are all functional herein. It is also contemplated that methods, which generate basic catalyst vapors in situ, such as the addition of water to amine salts, or the addition of water to a silazane, such as hexamethyldisilazane, will also be effective.

- 5       The basic catalyst used may be at any concentration desired. For example, the concentration may be from about 1 ppm up to a saturated atmosphere.

The exposure can be at any temperature desired from room temperature up to about 300°C. A temperature in the range of from about 20°C to about 200°C is typical, with a range of from about 20°C to about 100°C being more typical.

- 10       The SiH resin coating should be exposed to the basic catalyst and water environment for the time necessary to hydrolyze the Si-H groups to form silanols (Si-OH) and for the silanols to at least partially condense to form Si-O-Si bonds. Generally, exposures of up to about 20 minutes are typical, with exposures of at least about 1 second up to about 5 minutes being more typical. If the coatings are to  
15 be used as a dielectric layer, it is generally typical to have a shorter exposure; as longer exposures tend to increase the dielectric constant of the coating.

- When the coating is exposed to the basic catalyst and water in the liquid state, the exposure is usually conducted by immersing the coated substrate in a solution. Other equivalent methods can be used, such as flushing the coating with a  
20 basic catalyst and water solution. In addition, vacuum infiltration may also be used to increase penetration of the basic catalyst and water into the coating.

- The basic catalyst solution used in this embodiment may be at any concentration desired. Generally when ammonium hydroxide is used, a concentrated aqueous solution of between about 28 and about 30% is typical since  
25 the duration of exposure is thereby shortened. When dilute solutions are used, the diluent is generally water.

- Exposure to the basic catalyst and water solution in this embodiment may be conducted at any temperature and pressure desired. Temperatures from about room temperature (20-30°C) up to about the boiling point of the basic catalyst  
30 solution, and pressures from below to above atmospheric are all contemplated herein. From a practical standpoint, it is typical that the exposure occur at about room temperature and at about atmospheric pressure.

The resin coating is exposed to the basic catalyst solution in this embodiment for the time necessary to hydrolyze the Si-H groups to form silanols (Si-OH) and for

the silanols to at least partially condense to form Si-O-Si bonds. Generally, exposures of up to about 2 hours are typical, with exposures of at least about 1 second up to about 15 minutes being more typical.

Alternatively, the coating may be exposed to both a liquid basic catalyst and water environment (e.g., ammonium hydroxide) and a gaseous basic catalyst and water vapor environment (ammonia gas and water vapor). The exposures may be either sequential or simultaneous, and are generally under the same conditions as those described above.

After the resin is exposed to one of the above environments, the solvent is then removed from the coating. This can be accomplished by any desired means, including but not limited to, heating the coating, and by vacuum. When the solvent is removed by heating the coating, condensation of the remaining silanols may be facilitated.

The coating produced by this process can be used as the starting material ("porous network coating") in the present invention. In a typical procedure to produce a porous network coating, a substrate is coated with the Si-H containing resin and solvent in a manner which ensures that at least about 5 volume % of the solvent remains in the coating. The coating is then exposed to the basic catalyst and water, and the solvent is evaporated.

Another method of making such a porous network coating is to thermally cure a siloxane resin containing large alkyl groups and to thermally decompose the alkyl groups to create porosity in the coating. As disclosed in U.S. Patent Nos. 6,143,360 and 6,184,260, to Zhong, which are hereby incorporated herein by reference, hydrosilicon containing resin was allowed to contact with a 1-alkene comprising about 8 to about 28 carbon atoms in the presence of a platinum group metal-containing hydrosilation catalyst, effecting formation of an alkylhydrosiloxane resin where at least about 5 percent of the silicon atoms are substituted with at least one hydrogen atom, and the resulting resin was heated at a temperature sufficient to effect curing of the resin and thermolysis of alkyl groups from the silicon atoms, thereby forming a nanoporous silicone resin.

U.S. Patent No. 6,232,424 and U.S. Patent Application Serial Nos. 425,306, 425,901, and 459,331, to Zhong et al., which are hereby incorporated herein by reference, disclose silicone resins and porous coatings made from the silicone resins. The silicone resins are made from a mixture comprising 15 to 70 mol % of

- tetraalkoxysilane, 12 to 60 mol % of an organosilane described by formula  $R'SiX_3$ , where  $R'$  is an hydrogen or alkyl group containing 1 to 6 carbon atoms, and 15 to 70 mol % of an organotrialkoxysilane described by formula  $R''Si(OR'')_3$ , where  $R''$  is a hydrocarbon group comprising about 8 to 24 carbon atoms or a substituted hydrocarbon group comprising a hydrocarbon chain having about 8 to 24 carbon atoms.

- U.S. Patent Application entitled "SILICONE RESINS AND POROUS MATERIALS PRODUCED THEREFROM", to Zhong, filed September 12, 2001 and hereby incorporated herein by reference, discloses porous coatings made from silicone resins having the general formula  $(R^1SiO_{3/2})_x(HSiO_{3/2})_y$  where  $R^1$  is an alkyl group having 8 to 24 carbon atoms. The coatings produced therein have a dielectric constant between 1.5 and 2.3. The above-referenced patent application further provides the following description of a porous low-k dielectric coating made in two steps from a resin with a formula of  $(R^1SiO_{3/2})_x(HSiO_{3/2})_y$  where R is 3,7,11,15-tetramethyl-3-hydroxy-hexadecyl.

- U.S. Patent Application entitled "SILICONE RESINS AND POROUS MATERIALS PRODUCED THEREFROM", to Zhong, filed September 12, 2001 and hereby incorporated herein by reference, discloses porous coatings made from silicone resins having the general formula  $(R^1SiO_{3/2})_u(HSiO_{3/2})_v(SiO_{4/2})_w(HOSiO_{3/2})_z$  where  $R^1$  is a branched alkyl group having 8 to 24 carbon atoms containing at least one electron-withdrawing group in a pendant position on the alkyl chain; u has a value of 0.1 to 0.7; v has a value of 0.12 to 0.6;  $z \geq 0.5$ ;  $w + z$  has a value of 0.15 to 0.7; and  $u + v + w + z = 1$ .

- Step 1. A resin sample was prepared by combining components (A), (B), (C), (D), (E), and (F) as described below in the amounts described in Table 1 of the above-referenced U.S. patent application:

- (A) 0.45 mole parts of triethoxysilane,  
 (B) 0.25 mole parts of an organotriethoxysilane,  $RSi(OR')_3$   
 where R is 3,7,11,15-tetramethyl-3-hydroxy-hexadecyl,  
 (C) 0.30 mole parts of tetraethoxysilane, and

(D) a mixture of methyl isobutyl ketone (MIBK) and isobutyl isobutyrate (6:4 weight ratio), enough to make the concentration of the resulting resin 9%.

To this mixture was added a mixture of (E) water and (F) hydrogen chloride in the amounts described in Table 1 of the above-referenced application. The resulting reaction product was stripped of volatiles under reduced pressure at 60 °C until the solid content became 14 to 21%. Isobutyl isobutyrate was added to make the solid content 14%. The solution was then heated to reflux for 2 hours and water produced was removed continuously. The solvent was then changed to cyclohexanone by stripping off isobutyl isobutyrate and adding cyclohexanone.

Step 2. The resulting resin solution was spin-coated onto silicon wafers suitable for dielectric constant measurements, and cured in a nitrogen flow at 440°C for 1 hour. The dielectric constant was measured as 1.9. Alternatively, the curing of the spin-coated films may be accelerated with plasma and/or UV-assisted processes.

U.S. Patent Application Serial No. 915,899, which is hereby incorporated herein by reference, discloses porous coatings from resins containing  $(\text{RSiO}_{3/2})(\text{R}'\text{SiO}_{3/2})(\text{R}''\text{SiO}_{3/2})$  resins wherein R is an alkyl group having 1 to 5 carbon atoms or a hydrogen atom, R' is a branched alkoxy group and R'' is a substituted or un-substituted linear, branched, or cyclic monovalent organic group having 6 to 30 carbon atoms.

U.S. Patent Application Serial Nos. 915,903 and 915,902, which are hereby incorporated herein by reference, disclose porous coatings made from resins of the formula  $\text{TRTR}'$  where R is either a methyl or hydrogen group and a R' is a branched alkoxy group.

Although porous dielectric materials having low dielectric constants are desirable, it would be advantageous to have a porous dielectric material with a higher elastic modulus.

In order to raise the elastic modulus of the porous dielectric material, it is exposed to a UV cure. The UV curing process improves the mechanical properties of the porous low-k dielectric material, increasing material hardness while maintaining the dielectric pore, structure, density, and electrical properties.

- 5 In a typical UV curing process, a UV radiator tool is utilized, which is first purged with nitrogen or argon to allow the UV radiation to enter the process chamber with minimal spectral absorption. The process chamber is purged separately and process gases, such as O<sub>2</sub>, N<sub>2</sub>, H<sub>2</sub>, Ar, He, C<sub>x</sub>H<sub>y</sub>, air, and mixtures thereof, may be utilized for different applications. UV generating bulbs with different spectral
- 10 distributions may be selected depending on the application. The wafer temperature may be controlled ranging from room temperature to 450°C, and the process pressure can be less than, greater than, or equal to atmospheric pressure.

Examples of typical UV cure conditions for a 200 mm wafer are shown below.

15	UV Power:	0 mW - 1000 mW/cm <sup>2</sup>
	UV wavelength:	continued spectral distribution from 100-600nm
	Wafer Temperature:	room temp. - 450°C
	Process Pressure:	<, >, or = to atmospheric
	UV Cure Time:	< 300 seconds
20	Plasma Gases:	H <sub>2</sub> /N <sub>2</sub> /C <sub>x</sub> H <sub>y</sub> /O <sub>2</sub> :
	Forming Gas (FG) Flow Rate:	purge
	O <sub>2</sub> Flow Rate:	purge
	N <sub>2</sub> Flow Rate:	purge
25	H <sub>2</sub> /N <sub>2</sub> Gas Mixture flow rate:	purge

- The elastic modulus of the UV cured porous dielectric materials is increased as compared to a furnace (thermally) cured porous dielectric material, which would have an elastic modulus of between about 1.0 GPa and about 3.5 GPa when the dielectric constant is between about 1.6 and about 2.4. This increase in the elastic
- 30 modulus is typically greater than about 50%. Typically, the elastic modulus of the UV cured porous dielectric material is greater than about 2.5 GPa, and more typically between about 4 GPa and about 10 GPa.

- The UV cured porous dielectric materials of the present invention have improved chemical stability and improved dimensional stability. By improved
- 35 chemical stability, we mean that the porous dielectric materials are more resistant to chemicals, such as cleaning solutions and chemical polishing solutions, and plasma damaging during photoresist ashing and dry etching processes.



However, UV cure can generate a notable amount of polar species in the porous dielectric materials.

- The UV cured porous dielectric materials can optionally be post-UV treated using any type of thermal and/or plasma exposure to reduce the dielectric constant, if desired. For example, the UV cured porous dielectric materials can be annealed by placing the materials in a conventional oven until the polar species are removed, such as at a temperature of between about 400°C and about 450°C for between about 30 and about 60 minutes. An alternative process for annealing the materials involves annealing the UV cured porous dielectric materials in a Rapid Anneal Processing (RAP) chamber in order to reduce the dielectric constant. The UV cured porous dielectric material is annealed at a typical temperature for a sufficient time, and cooled to about 100°C. However, RAP may not be necessary in some applications.

Typical operating conditions for the RAP process are shown below.

- |    |                    |                |
|----|--------------------|----------------|
| 15 | Ramp rate:         | 15 - 150°C/sec |
|    | Wafer Temperature: | 150 - 450°C    |
|    | Annealing Time:    | < 120 seconds  |
|    | Process Pressure:  | atmospheric    |

- A third type of post-UV treatment that can be used involves the exposure of the UV cured porous dielectric materials to a plasma condition at elevated temperatures. In a typical plasma-assisted post-UV treatment, process gases, such as O<sub>2</sub>, N<sub>2</sub>, H<sub>2</sub>, Ar, He, C<sub>x</sub>H<sub>y</sub>, fluorine-containing gas, and mixtures thereof, may be utilized for different applications. The wafer temperature may be controlled ranging from room temperature to 450°C. Typically, the UV cured porous dielectric material is plasma treated at a process pressure between about 1 Torr and about 10 Torr.

Examples of typical plasma-assisted post-UV treatment conditions for 200 mm and 300 mm wafers are shown below.

- | 30 | <b>Condition</b>                         | <b>200 mm system</b>   | <b>300 mm system</b> |
|----|--|--|----------------------|
|    | Microwave Plasma Power:                  | 500 W – 3000 W   | 500 W – 3000 W       |
|    | Wafer Temperature:                       | 80°C - 350°C   | 80°C - 350°C         |
|    | Process Pressure:                        | 1.0 Torr – 3.0 Torr  | 1.0 Torr – 4.0 Torr  |
|    | Plasma Treatment Time:                   | < 90 seconds   | < 90 seconds         |
| 35 | Plasma Gases:                            | H <sub>2</sub> /N <sub>2</sub> /CF <sub>4</sub> /O <sub>2</sub> /Ar/He/C <sub>x</sub> H <sub>y</sub> |                      |
|    | N <sub>2</sub> H <sub>2</sub> Flow Rate: | > 0 – 4000 sccm  | > 0 – 10,000 sccm    |

O <sub>2</sub> Flow Rate:	> 0 – 4000 sccm	> 0 – 10,000 sccm
CF <sub>4</sub> Flow Rate:	> 0 – 400 sccm	> 0 – 1000 sccm
Ar Flow Rate:	> 0 – 4000 sccm	> 0 – 10,000 sccm
He Flow Rate:	> 0 – 4000 sccm	> 0 – 10,000 sccm

The dielectric constant of the post-UV treated, UV cured porous dielectric materials is reduced as compared to the UV cured porous dielectric materials. The dielectric constant of the post-UV treated, UV cured porous dielectric materials is typically between about 1.1 and about 3.5 and more typically between about 1.6 and about 2.4.

Typical material properties of porous low-k films with UV curing are shown in Table 1 below.

**Table 1. Porous Low-K Material Characteristics with UV Curing**

Material Properties	MSQ-Based Porous Dielectric Material	HSQ-Based Porous Dielectric Material
Change in Dielectric Constant	< 0.1	< 0.2
Modulus Increase	> 50%	>50%
Porosity	Unchanged	Unchanged
Moisture Absorption	Hydrophobic	Hydrophilic
Plasma Curing Chemistry	O <sub>2</sub> N <sub>2</sub> /H <sub>2</sub>	N <sub>2</sub> /H <sub>2</sub> O <sub>2</sub>
UV Curing Purge gases	O <sub>2</sub> , Ar, He, air, N <sub>2</sub> /H <sub>2</sub>	N <sub>2</sub> /H <sub>2</sub> , Ar, He, air, O <sub>2</sub>
Density	Unchanged	Unchanged
Thickness Loss	< 10%	< 10%
Refractive Index Change	< 0.01	< 0.03

In order that the invention may be more readily understood, reference is made to the following examples, which are intended to illustrate the invention, but not limit the scope thereof.

The following graphs show the attributes of the UV curing. Examples are presented for (i) blanket MSQ-based porous low-k thin films, (ii) blanket MSQ-based porous low-k thin films, and (iii) blanket MSQ/HSQ mixed porous low-k films, all with a thickness of approximately 5000 Å.

#### Example 1: MSQ-based porous low-k film

The FTIR spectra of MSQ-based porous low-k films are characterized by Si-CH<sub>3</sub> characteristic bands near 1280 cm<sup>-1</sup> and 3000 cm<sup>-1</sup>, as well the two Si-O peaks

near  $1100\text{ cm}^{-1}$ . The curing of these films which typically results in a modulus increase of 100% or more is reflected in the FTIR spectra by the partial removal of the Si-CH<sub>3</sub> bands, and the change in the ratio of cage (higher wavenumber) vs. network (lower wavenumber) Si-O bonds.

- 5        The UV treatment can successfully cure the MSQ-based porous low-k films. Fig. 1 shows FTIR spectra for MSQ-based porous low-k films: thermally cured (bottom), thermally + UV cured in O<sub>2</sub> for 1 minute (center), and thermally + UV cured in N<sub>2</sub> for 5 minutes (top). The curing effectiveness is strongly dependent on the chamber purge gas composition. It has been observed that O<sub>2</sub> is more effective for
- 10    the UV curing than N<sub>2</sub>.

Example 2: HSQ-based porous low-k film (Dow Corning's XLK)

- 15        The FTIR spectra of HSQ-based porous low-k films are characterized by the Si-H band near  $2200\text{ cm}^{-1}$ , a band near  $850\text{ cm}^{-1}$  which is attributed to a SiO-H stretch mode and the two Si-O peaks near  $1100\text{ cm}^{-1}$ . The curing of these films which typically results in a modulus increase of 100% or more is reflected in the FTIR spectra by the complete removal of the Si-H band, and the change in ratio of cage vs. network Si-O bonds.

- 20        The UV treatment can successfully cure the HSQ-based porous low-k films. Fig. 2 shows FTIR spectra for HSQ-based porous low-k films: uncured (green), UV cured for 60 seconds in O<sub>2</sub> (purple) and N<sub>2</sub> (blue). However, the efficiency for the curing is dependent on the chamber purge gas composition. It has been observed that O<sub>2</sub> is more effective for the UV curing than N<sub>2</sub>.

25    Example 3: HSQ/MSQ mixed porous low-k film (Dow Corning)

- 30        The FTIR spectra of HSQ/MSQ-mixed porous low-k films are characterized by the usual Si-H band near  $2200\text{ cm}^{-1}$ , the SiO-H stretch mode band near  $850\text{ cm}^{-1}$  and the two Si-O peaks near  $1100\text{ cm}^{-1}$ . In addition there is the Si-CH<sub>3</sub> characteristic feature near  $1280\text{ cm}^{-1}$ . For this specific example a 5% MSQ/95% HSQ mixed film has been studied. The UV curing of these films exhibits a much stronger dependence on the UV purge gas mixture than the pure HSQ-based porous film. Nevertheless, the O<sub>2</sub> purged UV treatment results in an effective and successful curing of the low-k films.

Fig. 3 shows FTIR spectra for 5% MSQ/95% HSQ-based porous low-k films: uncured (blue), UV cured for 60 seconds in O<sub>2</sub> (purple), and N<sub>2</sub> (green). For all cases a subsequent or possibly concomitant anneal step is necessary in order to remove the Si-OH bonds which are typically generated during the UV curing process.

While certain representative embodiments and details have been shown for purposes of illustrating the invention, it will be apparent to those skilled in the art that various changes in the compositions and methods disclosed herein may be made without departing from the scope of the invention, which is defined in the appended claims.

What is claimed is:

## CLAIMS

1. A process for making a UV cured material having improved properties comprising:
  - 5 providing a porous dielectric material having a first dielectric constant and having a first elastic modulus; and
  - UV curing the porous dielectric material to produce a UV cured porous dielectric material having a second dielectric constant which is comparable to the first dielectric constant and having a second elastic modulus which is greater than  
10 the first elastic modulus.
2. The process of claim 1 wherein the porous dielectric material is selected from a hydrogen silsesquioxane dielectric material, a methylsilsesquioxane dielectric material, an organic dielectric material, an inorganic dielectric material, or a  
15 combination thereof.
3. The process of claim 1 wherein the porous dielectric material is produced by a spin-on process or a chemical vapor deposition process.
- 20 4. The process of claim 1 wherein the porous dielectric material is selected from a porogen-generated porous dielectric material, a solvent-based porous dielectric material, or a molecular engineered porous dielectric material, or combinations thereof.
- 25 5. The process of claim 1 wherein the porous dielectric material is UV cured for no more than about 300 seconds.
6. The process of claim 1 wherein the porous dielectric material has a wafer temperature that is less than about 450°C during UV curing.
- 30 7. The process of claim 1 wherein the porous dielectric material has a wafer temperature that is between about room temperature and about 450°C during UV curing.

8. The process of claim 1 wherein the porous dielectric material is UV cured at a process pressure that is less than atmospheric pressure, greater than atmospheric pressure, or equal to atmospheric pressure.
- 5 9. The process of claim 1 wherein the porous dielectric material is UV cured at a UV power between about 0 and about 1000 mW/cm<sup>2</sup>.
10. The process of claim 1 wherein the porous dielectric material is UV cured with a gas purge, wherein the gas is selected from the group consisting of N<sub>2</sub>, O<sub>2</sub>, Ar, He, H<sub>2</sub>, C<sub>x</sub>H<sub>y</sub>, air, and combinations thereof.
- 10 11. The process of claim 1 wherein the porous dielectric material is UV cured using a UV wavelength spectrum between about 100nm and about 400nm.
- 15 12. The process of claim 1 wherein the increase in elastic modulus between the first elastic modulus of the porous dielectric material and the second elastic modulus of the UV cured porous dielectric material is greater than about 50%.
13. The process of claim 1 wherein the second elastic modulus of the UV cured porous dielectric material is greater than about 2.5 GPa.
- 20 14. The process of claim 1 wherein the second elastic modulus of the UV cured porous dielectric material is between about 4 GPa and about 10 GPa.
- 25 15. The process of claim 1 further comprising post-UV treating the UV cured porous dielectric material to provide a post-UV treated, UV cured porous dielectric material having a third dielectric constant which is less than the second dielectric constant and having a third elastic modulus which is comparable to the second elastic modulus.
- 30 16. The process of claim 15 wherein the third dielectric constant of the post-UV treated, UV cured porous dielectric material is between about 1.1 and about 3.5.
17. The process of claim 15 wherein the third dielectric constant of the post-UV treated, UV cured porous dielectric material is between about 1.6 and about 2.4.

18. The process of claim 15 wherein the post-UV treating is annealing.
19. The process of claim 18 wherein the UV cured porous dielectric material is  
5 annealed at a temperature less than about 450°C.
20. The process of claim 18 wherein the UV cured porous dielectric material is  
annealed at a temperature between about 150°C and about 450°C.
- 10 21. The process of claim 18 wherein the UV cured porous dielectric material is  
annealed for no more than about 60 minutes.
22. The process of claim 15 wherein the post-UV treating is plasma treating by  
exposing the UV cured porous dielectric material to a plasma condition at elevated  
15 temperatures.
23. The process of claim 22 wherein the UV cured porous dielectric material is  
plasma treated at a plasma power between about 500 W and about 3000 W.
- 20 24. The process of claim 22 wherein the UV cured porous dielectric material is  
plasma treated at a temperature between about 100°C and about 450°C.
25. The process of claim 22 wherein the UV cured porous dielectric material is  
plasma treated for no more than about 90 seconds.
- 25 26. The process of claim 22 wherein the UV cured porous dielectric material is  
plasma treated at a process pressure between about 1 Torr and about 10 Torr.
- 30 27. The process of claim 22 wherein the UV cured porous dielectric material is  
plasma treated with a plasma gas, wherein the plasma gas is selected from the  
group consisting of N<sub>2</sub>, O<sub>2</sub>, Ar, He, H<sub>2</sub>, C<sub>x</sub>H<sub>y</sub>, fluorine-containing gas, and  
combinations thereof.

28. A UV cured porous dielectric material prepared by the process of claim 1.
29. A post-UV treated, UV cured porous dielectric material prepared by the process of claim 15.
- 5 30. An electronic device containing a UV cured porous dielectric material prepared by the process of claim 1.
31. An electronic device containing a post-UV treated, UV cured porous dielectric material prepared by the process of claim 15.
- 10 32. A substrate having a UV cured coating prepared by the process of claim 1.
33. A substrate having a post-UV treated, UV cured coating prepared by the process of claim 15.
- 15 34. A UV cured porous dielectric material having a dielectric constant between about 1.1 and about 3.5 and an elastic modulus that is about 50% greater than a non-UV cured porous dielectric material.
- 20 35. A UV cured porous dielectric material having a dielectric constant between about 2.0 and about 2.9 and an elastic modulus that is about 50% greater than a non-UV cured porous dielectric material.



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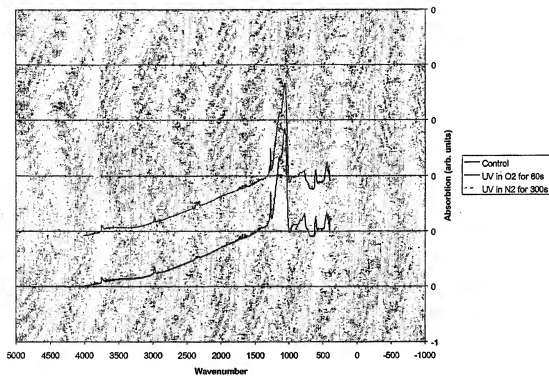


Fig. 1

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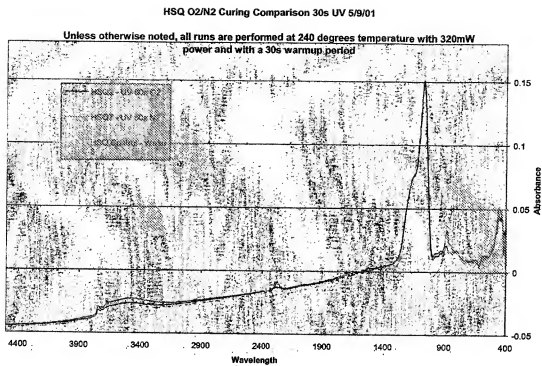


Fig. 2

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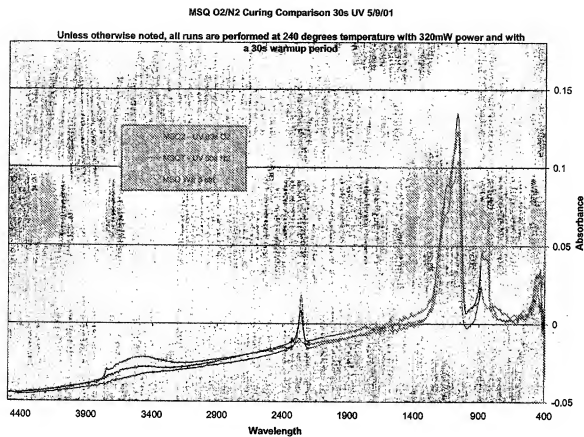


Fig. 3

# INTERNATIONAL SEARCH REPORT

International Application No.  
PCT/US 02/29173

<b>A. CLASSIFICATION OF SUBJECT MATTER</b> IPC 7 H01L21/312 H01L21/312 C08J9/00 C09D183/04 C09D183/02		
According to International Patent Classification (IPC) or to both national classification and IPC		
<b>B. FIELDS SEARCHED</b> Minimum documentation searched (classification system followed by classification symbols) IPC 7 H01L C09D C08J		
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched		
Electronic data base consulted during the international search (name of data base and, where practical, search terms used) EPO-Internal, WPI Data, PAJ		
<b>C. DOCUMENTS CONSIDERED TO BE RELEVANT</b>		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
P, X	US 2002/102413 A1 (JANG JEFF ET AL) 1 August 2002 (2002-08-01) cited in the application claims 1-29	1-4
P, X	page 2, paragraph 19 -page 3, paragraph 23 page 3, paragraph 33 -page 4, paragraph 43 & WO 01 70628 A (DOW CORNING CORPORATION) 27 September 2001 (2001-09-27) cited in the application	1-4
P, X	WO 01 86709 A (HONEYWELL INT INC) 15 November 2001 (2001-11-15) claims 1-35 page 18, line 28 -page 19, line 5 ----- -/-	1
<input checked="" type="checkbox"/> Further documents are listed in the continuation of box C. <input checked="" type="checkbox"/> Patent family members are listed in annex.		
* Special categories of cited documents : "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "Z" document member of the same patent family		
Date of the actual completion of the international search  7 January 2003		Date of mailing of the international search report  15/01/2003
Name and mailing address of the ISA European Patent Office, P.B. 5618 Patentlaan 2 NL - 2280 HV Rijswijk Tel. (+31-70) 340-2040, Tx. 31 651 epo nl, Fax: (+31-70) 340-3016		Authorized officer  Depijper, R

## INTERNATIONAL SEARCH REPORT

 International Application No.  
 PCT/US 02/29173

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 1 035 183 A (CATALYSTS & CHEM IND CO) 13 September 2000 (2000-09-13) claims 1,3,7,8 page 7, line 51 - line 55 -----	1
A	EP 0 962 965 A (EATON CORP) 8 December 1999 (1999-12-08) claims 1-23 column 3, line 33 -column 6, line 27 -----	1
A	EP 0 816 925 B (DOW CORNING ASIA LTD.) 26 September 2001 (2001-09-26) page 2, paragraph 19 - paragraph 23; claims 1-9 page 1, line 3 - line 7 -----	1
A	EP 0 462 715 B (DOW CORNING) 27 December 1991 (1991-12-27) claim 1 page 2, line 51 -page 3, line 18 -----	1
A	US 4 900 582 A (NAKAYAMA MUNEO ET AL) 13 February 1990 (1990-02-13) claims 1-22 -----	1
A	EP 1 108 763 A (DOW CORNING) 20 June 2001 (2001-06-20) claim 1 page 4, line 46 -page 5, line 20 -----	1

## INTERNATIONAL SEARCH REPORT

Information on patent family members

Information on application No.

PCT/US 02/29173

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US 2002102413 A1	01-08-2002	US 2001038919 A1 AU 4590501 A EP 1265813 A2 WO 0170628 A2 US 2002106500 A1	08-11-2001 03-10-2001 18-12-2002 27-09-2001 08-08-2002
WO 0186709 A	15-11-2001	AU 6117501 A WO 0186709 A2	20-11-2001 15-11-2001
EP 1035183 A	13-09-2000	EP 1035183 A1 US 6451436 B1 WO 0018847 A1	13-09-2000 17-09-2002 06-04-2000
EP 0962965 A	08-12-1999	US 6417115 B1 EP 0962965 A2 JP 2000036494 A	09-07-2002 08-12-1999 02-02-2000
EP 0816925 B	07-01-1998	DE 69706920 D1 DE 69706920 T2 EP 0816925 A1 JP 10083080 A US 5861235 A	31-10-2001 25-04-2002 07-01-1998 31-03-1998 19-01-1999
EP 0462715 B	27-12-1991	US 5059448 A CA 2043411 A1 DE 69117353 D1 DE 69117353 T2 EP 0462715 A1 ES 2086485 T3 JP 2561979 B2 JP 4298542 A KR 9501671 B1	22-10-1991 19-12-1991 04-04-1996 05-09-1996 27-12-1991 01-07-1996 11-12-1996 22-10-1992 28-02-1995
US 4900582 A	13-02-1990	JP 2624254 B2 JP 63289939 A	25-06-1997 28-11-1988
EP 1108763 A	20-06-2001	US 6143360 A EP 1108763 A2 JP 2001240673 A SG 89363 A1 US 6184260 B1	07-11-2000 20-06-2001 04-09-2001 18-06-2002 06-02-2001